IONIC BONDING

2D

3D

- Electrons are transferred from metals to non-metals creating full outer shells
- Atoms become ions

H-H Single

O=**O** Double

N = N Triple

- Componds form giant lattices
- Electrostatic attraction between oppositely charged ions creates a strong bond



ALL & STICK

Positive

metal ion

Delocalised

electron

CENTURY

+

- Outer electrons are shared between two non-metal atoms so each atom has a full outer shell
- Each **shared pair** of electrons creates one covalent bond
- Sharing electrons creates a strong bond



METALLIC BONDING

2D

3D

- Giant, lattice structures
- Atoms lose outer electrons becoming positive metal ions
- Electrons are delocalised
- Electrostatic attraction between positive ions & negative electrons creates a strong bond